

Zirconium Encyclopedia Article

Zirconium

The following sections of this BookRags Literature Study Guide is offprint from Gale's For Students Series: Presenting Analysis, Context, and Criticism on Commonly Studied Works: Introduction, Author Biography, Plot Summary, Characters, Themes, Style, Historical Context, Critical Overview, Criticism and Critical Essays, Media Adaptations, Topics for Further Study, Compare & Contrast, What Do I Read Next?, For Further Study, and Sources.

(c)1998-2002; (c)2002 by Gale. Gale is an imprint of The Gale Group, Inc., a division of Thomson Learning, Inc. Gale and Design and Thomson Learning are trademarks used herein under license.

The following sections, if they exist, are offprint from Beacham's Encyclopedia of Popular Fiction: "Social Concerns", "Thematic Overview", "Techniques", "Literary Precedents", "Key Questions", "Related Titles", "Adaptations", "Related Web Sites". (c)1994-2005, by Walton Beacham.

The following sections, if they exist, are offprint from Beacham's Guide to Literature for Young Adults: "About the Author", "Overview", "Setting", "Literary Qualities", "Social Sensitivity", "Topics for Discussion", "Ideas for Reports and Papers". (c)1994-2005, by Walton Beacham.

All other sections in this Literature Study Guide are owned and copyrighted by BookRags, Inc.

Contents

| | |
|---|-------------------|
| Zirconium Encyclopedia Article..... | 1 |
| Contents..... | 2 |
| Zirconium..... | 3 |

Zirconium

Zirconium is the second element in Group 4 of the **periodic table**, one of the transition **metals**. It has an **atomic number** of 40, an atomic **mass** of 91.22, and a chemical symbol of Zr.

Properties

Zirconium usually occurs as a hard, grayish white metal, whose surface has a flaky appearance. It may also be prepared as a black or bluish black powder. Zirconium's melting point is 3,375°F (1,857°C), its **boiling point** is 6,471°F (3,577°C), and its **density** is 6.5 grams per cubic centimeter. Zirconium is a fairly inactive element that combines slowly with **oxygen** in the air, but does not react with most cold acids or with **water**. One of its most important and useful properties is its transparency to neutrons. Most metals tend to absorb neutrons that pass through them, but zirconium is largely transparent to them.

Occurrence and Extraction

Zirconium is a relatively common element in the Earth's crust with an abundance estimated at about 150-230 parts per million. Its most common ores are zircon (zirconium silicate; $ZrSiO_4$) and baddeleyite (zirconia or zirconium **oxide**; ZrO_2). The largest suppliers of zirconium in the world are South Africa and Australia.

Discovery and Naming

Zirconium was discovered in 1789 by the German chemist **Martin Heinrich Klaproth**. Klaproth found the element in a stone brought to him from the island of Ceylon (now Sri Lanka). The stone was commonly known as a jacinth (or hyacinth) stone and had been known for many years. In fact, Saint John talks about the jacinth stone as one of the jewels found in the walls surrounding Jerusalem. Klaproth chose zirconium as a name for the new element from the Persian name for the same stone, *zargun*, meaning "gold-like."

Uses

A very small amount of zirconium is used in the manufacture of alloys for products such as flash bulbs, rayon spinnerets, lamp filaments, precision tools, and surgical instruments. About 95% of all the element produced, however, is converted to compounds, the two most important of which are zircon (zirconium silicate) and zirconia (zirconium oxide). Synthetic zircon is used to make gemstones that resemble fine diamonds and as a refractory material in foundry molds and furnace linings. Zircon and zirconium are both used extensively as abrasives in industrial operations.