

Acetone Encyclopedia Article

Acetone

The following sections of this BookRags Literature Study Guide is offprint from Gale's For Students Series: Presenting Analysis, Context, and Criticism on Commonly Studied Works: Introduction, Author Biography, Plot Summary, Characters, Themes, Style, Historical Context, Critical Overview, Criticism and Critical Essays, Media Adaptations, Topics for Further Study, Compare & Contrast, What Do I Read Next?, For Further Study, and Sources.

(c)1998-2002; (c)2002 by Gale. Gale is an imprint of The Gale Group, Inc., a division of Thomson Learning, Inc. Gale and Design and Thomson Learning are trademarks used herein under license.

The following sections, if they exist, are offprint from Beacham's Encyclopedia of Popular Fiction: "Social Concerns", "Thematic Overview", "Techniques", "Literary Precedents", "Key Questions", "Related Titles", "Adaptations", "Related Web Sites". (c)1994-2005, by Walton Beacham.

The following sections, if they exist, are offprint from Beacham's Guide to Literature for Young Adults: "About the Author", "Overview", "Setting", "Literary Qualities", "Social Sensitivity", "Topics for Discussion", "Ideas for Reports and Papers". (c)1994-2005, by Walton Beacham.

All other sections in this Literature Study Guide are owned and copyrighted by BookRags, Inc.

Contents

Acetone Encyclopedia Article.....	1
Contents.....	2
Acetone.....	3

Acetone

Acetone (C_3H_6O) is a colorless liquid that is used as a solvent in products, such as in nail polish and paint, and in the manufacture of other **chemicals** such as **plastics** and fibers. It is a naturally occurring compound that is found in plants and is released during the **metabolism** of fat in the body. It is also found in volcanic gases, and is manufactured by the chemical industry. Acetone is also found in the **atmosphere** as an oxidation product of both natural and **anthropogenic** volatile organic compounds (VOCs). It has a strong smell and taste, and is soluble in water. The evaporation point of acetone is quite low compared to water, and the chemical is highly flammable. Because it is so volatile, the acetone manufacturing process results in a large percentage of the compound entering the atmosphere. Ingesting acetone can cause damage to the tissues in the mouth and can lead to unconsciousness. Breathing acetone can cause irritation of the eyes, nose, and throat; headaches; dizziness; nausea; unconsciousness; and possible coma and death. Women may experience menstrual irregularity. There has been concern about the carcinogenic nature of acetone, but laboratory studies, and studies of humans who have been exposed to acetone in the course of their occupational activities show no evidence that acetone causes **cancer**.